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## **Acids Bases And Salts Answers**

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**NCERT Solutions for class 10 chapter 2 Acids, bases and**

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## **salt**

Answer-1. Post-Your-Explanation-1. 2. Bases react with acids to produce salts and water. water to produce acids and salts. salts to produce acids and water. neither acids, salts, nor water.

Answer-2.

## **Acids Bases and Salts Multiple Choice Questions and Answers**

A reaction in which an acid and base react with each other to give a salt and water is termed as neutralization reaction. For Example: (i)  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$  (ii)  $\text{HNO}_3 + \text{KOH} \rightarrow \text{KNO}_3 + \text{H}_2\text{O}$

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Title: Chemistry Workbook Answers Acids Bases And Salts

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Subject: Chemistry Workbook Answers Acids Bases And Salts

## **Chemistry Workbook Answers Acids Bases And Salts**

Short Answer type Question 10 Give a reason why (a) Great care should be taken while handling laboratory acids and bases. (b) Solid powder of baking soda does not turn dry litmus paper to blue in color. Question 11 Define acid rain and how it is produced? What are its harmful effect? Question 12 Define neutralization reaction with an example.

## **Worksheet for Class 7 Science Chapter 5: Acid Base and Salt**

Answer. (a) HCl is acid and NaOH is base whose combination forms the common salt. Its formula is NaCl (Sodium chloride). It is obtained from sea water. (b) Rock salt is the common name for the mineral "halite". Its chemical formula is NaCl. It may be white or light blue or yellow depending upon impurities present

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in it.

## **Acids Bases and Salts Chapter Wise Important Questions**

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Here both acids and bases get neutralised For example, when sodium hydroxide (NaOH) is added to hydrochloric acid (HCl), sodium chloride (NaCl) and water (H<sub>2</sub>O) are obtained.  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{Heat}$  6. Mark 'T' if the statement is true and 'F' if it is false:

## **NCERT Solutions Class 7 Science Chapter 5 Acid Base and**

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An acid is a substance that reacts with water to produce hydronium ions, H<sub>3</sub>O<sup>+</sup>. A base is a substance that produces hydroxide ions, OH<sup>-</sup>, in water. CHEM 0012 Lecture Notes 3. When an acid and base reacts, an ionic compound is produced as one of the products. We call ionic compounds "salts".

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## **Acids, Bases and Salts**

The salts of strong acids and weak bases give acidic solution having pH less than 7. Example,  $\text{NH}_4\text{Cl}$ , Ammonium Chloride will have pH less than 7. The salts of weak acids and strong bases give basic solution having pH more than 7. Example,  $\text{Na}_2\text{CO}_3$ , Sodium Carbonate will have pH more than 7. Question 7.

## **Acids, Bases and Salts Class 10 Important Questions with**

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neutralization reaction. a reaction in which an acid and a base react in an aqueous solution to produce a salt and water (19.4)  
pH. a number used to denote the hydrogen-ion concentration, or acidity, of a solution; it is the negative logarithm of the hydrogen-ion concentration of a solution (19.2) salt hydrolysis.

## **Chapter 19 Acids, Bases, and Salts Flashcards | Quizlet**

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Play this game to review Chemistry. Which of the following values would represent the pH of a strong base?

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## **Important Questions for CBSE Class 10 Science Chapter 2**

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## **Acids bases and salts//mcq//quiz answers//competitive ...**

question\_answer22) A base reacts with an acid to form salt and

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water. What is this reaction called? A) Oxidation reaction done clear. B) Neutralization reaction done clear. C) Reduction reaction done clear. D) Biochemical reaction done clear. View Solution play\_arrow; question\_answer23) Identify the base used in making antacids.

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