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A. Bonding Principles 1. Most bonds are polar covalent; the more electronegative atom is the “-” end of the bond
Example: For $>C=O$, O is negative, C is positive 2. Simplest Model: Lewis Structure: Assign valence electrons as bonding electrons and non-bonding lone-pairs; more accurate bonding models include valence- bonds, molecular orbitals and molecular modeling

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