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OXIDATION AND REDUCTION Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent. 1. $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ 2. $2\text{Li} + \text{S} \rightarrow 2\text{Li}_2\text{S}$ 3. $2\text{Cs} + \text{Br}_2 \rightarrow 2\text{CsBr}$ 4. $3\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$ 5. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$

05 CTR ch20 7/12/04 8:17 AM Page 517 THE MEANING OF ...

sentence that is true about oxidation a
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Oxidation-Reduction Reactions 221
SECTION 201 THE MEANING OF
OXIDATION AND REDUCTION (pages
631-638) This section explains oxidation
and reduction in terms of

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Reducing agent. reactions that involve the processes of oxidation ...

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Reactions Chapter 20 Notes Oxidation-Reduction Reactions 20.1 The Meaning of Oxidation and Reduction What are Oxidation and Reduction? o Oxygen and Redox KEY = The substance gaining O is oxidized, while the substance losing O is reduced Oxidation-Reduction Reactions = Reactions with a substance being Page 2/11

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General Chemistry: Principles and Modern Applications, Petrucci, Harwood, Herring, Madura, 9th edition, 2007, Pearson Prentice Hall, p. 159 Chapter 20: Electrochemistry; Introduction Practice Example: Balance the following equation in acidic solution: $\text{MnO}_4^- + \text{I}^- \rightarrow \text{Mn}^{2+} + \text{I}_2$ (s) Now we have an easier way to deal with this problem...

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Introduction In an oxidation-reduction
(or redox) reaction, electrons are
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