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One thing to note about equilibrium is that the reactions do not stop; both the forward reaction and the reverse reaction continue to occur. They both occur at the same rate, so any overall change by one reaction is cancelled by

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the reverse reaction. We say that chemical equilibrium is dynamic, rather than static. Also, because both reactions are occurring simultaneously, the equilibrium can be written backward.

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examples. Write equilibrium constant  
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is and how it shows whether the reaction favors the...

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system in equilibrium . Find the reaction . quotient,  $Q$ , ( $K_{eq}$ ) for a reaction equation . Solve for  $Q$  given . initial concentrations or pressures. Using . LeChatlier's principle, determine the direction of shift in an equilibrium when conditions are changed (concentration, pressure, phase change) Setup a .  $K_{sp}$  expression for a slightly soluble product. Calculate the

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o The concept of equilibrium stems from  
the study of kinetics. o In future units,  
the concept of equilibrium is essential to  
understand the nature of acids and  
bases and the study of insoluble salts. o  
The concept of equilibrium is essential in  
the study of biochemistry as well (O<sub>2</sub>  
transport; blood pH

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$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ . the ratio of  
the right side over the left if the system  
is not.... If  $Q < K_{eq}$ , the product side is  
too high and the equilibrium wil....  
Characteristics of equilibrium. occurs in  
reversible reactions, there is no nt



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change in conce.... equilibrium constant.  
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